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Chapter 12 States Of Matter

all matter is made of small particles, atoms and molecules, that are in constant motion
characteristics of a gas -all particles are in constant, random motion -all collisions between particles are perfectly elastic

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Chapter 12 States of Matter. kinetic energy. kinetic-molecular theory. elastic collision. temperature.
□□energy due to motion. □□□□□□describes the behavior of matter in terms of particles.... □□□□an elastic collision is one in which no kinetic energy is.... □□temperature is a measure of the average kinetic energy of th....

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Study Guide - Chapter 12 - States of Matter

46 terms. yiowin. Chapter 12 States of Matter. STUDY. PLAY. kinetic energy. □□energy due to motion. kinetic-molecular theory. □□□□□□describes the behavior of matter in terms of particles in motion, the model makes several assumptions about the size, motion, and energy of gas particles (page403)

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Chemistry: Matter and Change 7 Teacher Guide and Answers Chapter Assessment - Chapter 12 - States of Matter Reviewing Vocabulary 1. i 2. atom bonded to a small, highly electronegative atom e 3. d 4. j 5. a 6. f 7. b 8. k 9. h 10. c 11. g 12. s 13. n 14. p 15. r 16. l 17. o 18. Therefore, water will have a lower boiling point at an q 19. m

12 States of Matter

CHAPTER 12 SOLUTIONS MANUAL States of Matter Section 12.1 Gases pages 402-410 Practice Problems page 405 1. Calculate the ratio of effusion rates for nitrogen (N_2) and neon (Ne). Rate N_2 /Rate Ne 0.849 Rate N_2 / Rate Ne $\sqrt{\frac{\text{molar mass Ne}}{\text{molar mass } N_2}}$ 20.18 28.02 0.849 2.

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Calculate the ratio of diffusion rates for carbon monoxide and carbon dioxide. $1.25 \text{ Rate CO} = \text{Rate CO}_2 \sqrt{\quad}$

States of Matter

There are actually four well-known states of matter: solid, liquid, gas, and plasma. Plasma isn't represented in the iceberg photo, but the other three states of matter are. The iceberg itself consists of water in the solid state, and the lake consists of water in the liquid state.

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Section 12.1 Gases Section 12.2 Forces of Attraction Section 12.3 Liquids and Solids Section 12.4 Phase Changes Exit CHAPTER States of Matter 12 Click a hyperlink to view the corresponding slides.
•Use the kinetic-molecular theory to explain the

Chemistry: Matter and Change

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Pressure DEFINITION: Force per unit area Barometer - measures atmospheric (air) pressure
Manometer - measures gas in a closed container Air Pressure Units Dalton's Law of Partial Pressures
Dalton's Law of Partial Pressures Example SI Unit of pressure - Pascal (Pa) Air pressure

Chapter 12: States of Matter by Sydney Sturgeon

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Gases. Using the kinetic - molecular theory to explain the behavior of gases; describe how mass
affects the rates of diffusion and effusion; define the different ways to measure gas pressure;
calculate the partial pressure of a gas.

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Chapter 12 - States of Matter I. Kinetic Molecular Theory (KMT) and Gases A. KMT definition: matter
is composed of a bunch of tiny particles that are in constant motion; this results in a number of
phenomena that can be observed on the macroscopic level B. Gases 1.

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