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Chapter 4 Atomic Structure Workbook

Physical Science Reading and Study Workbook Level B Chapter 4 33 IPLS Chapter 4 Atomic Structure Summary 4.1 Studying Atoms • The ancient Greek philosopher Democritus believed that all matter consisted of extremely small particles that could not be divided. • Aristotle did not think there was a limit to the number of times matter could be ...

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Section 4.2 Structure of the Nuclear Atom One change to Dalton's atomic theory is that atoms are divisible into subatomic particles: Electrons, protons, and neutrons are examples of these fundamental particles There are many other types of particles, but we will study these three

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Chapter 4 Atomic Structure Electrons and Orbitals. Electrons and Orbitals. Use the table on page 117 of your textbook to find the ratio of the maximum number of electrons to the number of orbitals for each of four energy levels. 1.

Chapter 4 Atomic Structure Electrons and Orbitals

Chapter 4Atomic Structure Section 4.1 Studying Atoms (pages 100-105) This section discusses the development of atomic models. Reading Strategy(page 100) Summarizing As you read, complete the table about atomic models. For more information on this Reading Strategy, see the Reading and Study Skills in the Skills and Reference Handbook at the end of

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108 Chapter 4 108 Chapter 4 FOCUS Objectives 4.2.1 Identify three subatomic particles and compare their properties. 4.2.2 Distinguish the atomic number of an element from the mass number of an isotope, and use these numbers to describe the structure of atoms. Build Vocabulary Word-Part Analysis Have students look up the term isotope in a dictionary

4.2 The Structure of an Atom

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Chapter 4: Atomic Structure Vocabulary Review. the sum of the protons and neutrons in the nucleus of that atom. atoms of the same element that have different numbers of neutrons and different mass numbers. This activity was created by a Quia Web subscriber.

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