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Covalent Bonding
Single Covalent Bonds
When only one pair of
electrons is shared,
such as in a hydrogen
molecule, it is a single
covalent bond. The
shared electron pair is
often referred to as the
bonding pair. For a
hydrogen molecule,
shown in Figure 8.4,

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each covalently bonded atom equally attracts the pair of shared electrons.

Chapter 8: Covalent Bonding

covalent bonds. The majority of covalent bonds form between nonmetallic elements.

8. Describe how the octet rule applies to covalent bonds. Atoms share valence electrons; the shared electrons complete the

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octet of each atom. 9.

Illustrate the formation of single, double, and triple covalent bonds using Lewis structures.

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Bonding and Molecular

Structure 8-3 There are

two types of repulsive

forces between the two

atoms. First, the nuclei

repel because they are

both positively

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charged. Second, the electrons repel because they are both negatively charged. The attractive forces between the two atoms result from the

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8.2 The Nature of

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Experimental evidence, however, indicates ...

Chapter 8

1. In covalent bonds, electron sharing usually occurs so that atoms attain the electron configuration of noble gases. (8) 2. Atoms form double or triple covalent bonds if

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they can attain a noble gas structure by sharing two pairs or three pairs of electrons.

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vocabulary and key
questions for chapter 8
of the chemistry
textbook. I hope you
appreciate it as much
as I do, since it took
me 1 1/2 hours to fill
out this information. =)

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Gravity. Created by.
Petitt1. Terms in this
set (30) covalent bond.
a chemical bond that
involves sharing a pair
of electrons between
atoms in a molecule.
molecule. two or more
atoms held together by
covalent bonds.

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Bonding 8.1 Molecular
Compounds 1.

Important Definitions a.
Covalent Bond b.

molecule c. diatomic

molecules Our Special
Diatomic Elements H_2 ,

O_2 , N_2 , Cl_2 , Br_2 , I_2 , F_2

d. molecular compound
(in other words, it is
covalently bonded) e.

molecular formula 2.

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Bonding 71 SECTION 8.2 THE NATURE OF

COVALENT BONDING
(pages 217–220) This section uses electron dot structures to show the formation of single, double, and triple covalent bonds. It also describes and gives examples of coordinate covalent bonding, resonance structures, and exceptions to the octet rule.

SECTION 8.1 *Page 17/24*

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MOLECULAR COMPOUNDS (pages 213-216)

A covalent bond is formed between atoms held together by sharing electrons. A molecule is a group of atoms joined by covalent bonds.

Section 8.2

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When only one pair of

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such as in a hydrogen
molecule, it is a single,
11.

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universally compatible
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Covalent Bonds are formed when two nonmetals bond

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together. The elements carbon, oxygen, hydrogen, sulfur, nitrogen, phosphorus, chlorine, and bromine will be the main nonmetals studied in drawing dot structures of molecules. Bonds between these nonmetals are always Covalent. Part C of Chapter 3 should now be easy.

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Bonding Packet Answers Chapter 8 Concepts of Chemical Bonding .

Chemical Bonds Three
types: - Ionic

Electrostatic attraction
between ions Covalent

Sharing of electrons

Metallic Metal atoms

bonded to several

other atoms

Covalent Bonding • A

bond where electrons

from each atom are

shared

Chapter 8 Concepts of Chemical Bonding

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Chapter 8: Bonding 8.1
Lewis symbols and the
octet rule 8.2 Ionic
bonding 8.3 Covalent
bonding 8.4 Bond
polarity and
electronegativity 8.5
Drawing Lewis
structures 8.6
Resonance Structures
8.7 Exceptions to the
octet rule 8.8
Strengths and lengths
of covalent bonds

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